

Chapter 19

Practice SAT Chemistry Subject Test 3

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PRACTICE SAT CHEMISTRY SUBJECT TEST 3

You are about to take the third practice SAT Chemistry Subject Test. After answering questions 1–32, which constitute Part A, you'll be directed to answer questions 101–116, which constitute Part B. Then, begin again at question 33. Questions 33–69 constitute Part C.

When you're ready to score yourself, refer to the scoring instructions and answer key on [this page](#). Full explanations regarding the correct answers to all questions start on [this page](#).

MATERIAL IN THE FOLLOWING TABLE MAY BE USEFUL IN ANSWERING THE QUESTIONS IN THIS EXAMINATION.

PERIODIC TABLE OF THE ELEMENTS

1 H 1.0079																	2 He 4.0026
3 Li 6.941	4 Be 9.012											5 B 10.811	6 C 12.011	7 N 14.007	8 O 16.00	9 F 19.00	10 Ne 20.179
11 Na 22.99	12 Mg 24.30											13 Al 26.98	14 Si 28.09	15 P 30.974	16 S 32.06	17 Cl 35.453	18 Ar 39.948
19 K 39.10	20 Ca 40.48	21 Sc 44.96	22 Ti 47.90	23 V 50.94	24 Cr 52.00	25 Mn 54.938	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.39	31 Ga 69.72	32 Ge 72.59	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80
37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc (98)	44 Ru 101.1	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.75	52 Te 127.60	53 I 126.91	54 Xe 131.29
55 Cs 132.91	56 Ba 137.33	57 La 138.91	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm (145)	62 Sm 150.4	63 Eu 151.97	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.04	71 Lu 174.97	
87 Fr (223)	88 Ra 226.02	89 Ac 227.03	90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np 237.05	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (262)	
												§ Not yet named					

*Lanthanide Series	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm (145)	62 Sm 150.4	63 Eu 151.97	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.04	71 Lu 174.97
*Actinide Series	90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np 237.05	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (262)

Click [here](#) to see full page view

Note: For all questions involving solutions and/or chemical equations, assume that the system is in pure water unless otherwise stated.

Part A

Directions: Each set of lettered choices below refers to the numbered statements or questions immediately following it. Select the one lettered choice that best fits each statement or answers each question, and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–5 refer to the following.

- (A) Carbon
- (B) Nitrogen
- (C) Oxygen
- (D) Neon
- (E) Argon

1. Is the third most abundant gas in Earth's atmosphere
2. At standard conditions, has an allotropic form that is a good electrical conductor
3. Regardless of its electron configuration, it must always be paramagnetic when it's a single, neutrally charged atom
4. The key element delivered in soil fertilizer
5. Allotrope of this element is the primary absorber of UV solar radiation in Earth's atmosphere

Questions 6–9 refer to the following.

- (A) Chemical pH indicator
- (B) Acid/base buffer
- (C) Anhydrous solution
- (D) Hypotonic solution
- (E) Supersaturated solution

- 6. A conjugate acid/base pair with differing spectral absorbencies
- 7. An example of a solution not in equilibrium
- 8. Term used in reference to an aqueous solution's osmotic pressure
- 9. Addition of water to this solution will not change $[\text{H}_3\text{O}^+]$

Questions 10–14 refer to the following.

- (A) Standard voltaic potential
- (B) Entropy
- (C) Enthalpy
- (D) Reaction rate
- (E) Gibbs free energy

10. Increased with the addition of a catalyst

11. Abbreviated as “H”

12. A property that must decrease when a gas condenses into a liquid

13. Is always positive for a spontaneous chemical reaction

14. Is zero for a crystalline solid that is elementally pure at 0 K

Questions 15–19 refer to the following.

- (A) Alkali metals
- (B) Alkaline earth metals
- (C) Noble gases
- (D) Halogens
- (E) Transition metals

- 15. The most unreactive family of elements
- 16. Form negative ions in an ionic bond
- 17. Consist of atoms that have valence electrons in a *d* subshell
- 18. Exist as diatomic molecules at room temperature
- 19. Members possess the lowest first ionization energy in their respective period

Questions 20–24 refer to the following.

- (A) N_2
- (B) KI
- (C) CCl_4
- (D) AgNO_3
- (E) CaCO_3

- 20. A product of a neutralization of a strong acid with a strong base
- 21. A volatile covalent liquid at 25°C and 1 atm
- 22. Releases a gas with the addition of dilute acid
- 23. Forms a white precipitate when added to a solution of NaCl
- 24. Treatment of the dry solid with a mild oxidizing agent produces a purple solid

Questions 25–28 refer to the following.

- (A) Gamma decay
- (B) Nuclear fusion
- (C) Alpha decay
- (D) Positron emission
- (E) Nuclear fission

25. Is the principle reaction responsible for the energy output of the sun
26. Is a nuclear process that results in no change in the mass number and atomic number of a nuclide
27. Responsible for most helium found on Earth
28. The nuclear process that transmutes uranium-238 into thorium-234

Questions 29–32 refer to the following.

- (A) 0.1 *M* MgCl_2
- (B) 0.1 *M* HClO_4
- (C) 0.1 *M* NH_4OH
- (D) 0.1 *M* KOH
- (E) 0.1 *M* LiNO_3

- 29. Has a pH of 13
- 30. The solution with the lowest freezing point temperature
- 31. The solution with the highest boiling point temperature
- 32. Indicates a red flame when ionized with a Bunsen burner

PLEASE GO TO THE SPECIAL SECTION LABELED CHEMISTRY AT THE LOWER RIGHT-HAND CORNER OF THE ANSWER SHEET YOU ARE WORKING ON AND ANSWER QUESTIONS 101–116 ACCORDING TO THE FOLLOWING DIRECTIONS.

Part B

Directions: Each question below consists of two statements, I in the left-hand column and II in the right-hand column. For each question, determine whether statement I is true or false and whether statement II is true or false, and fill in the corresponding T or F ovals on your answer sheet. Fill in oval CE only if statement II is a correct explanation of statement I.

EXAMPLES:

I		II
EX 1. H_2SO_4 is a strong acid	BECAUSE	H_2SO_4 contains sulfur.
EX 2. An atom of oxygen is electrically neutral	BECAUSE	an oxygen atom contains an equal number of protons and electrons.

SAMPLE ANSWERS

	I	II	CE
EX 1	☐	☐	☐
EX 2	☐	☐	☐

I		II
101. Transition metal compounds are often colored	BECAUSE	they frequently possess partially filled <i>d</i> orbitals.
102. Chemical reactions slow down with lower temperature	BECAUSE	the energy barrier for the formation of products decreases with decreasing temperature.
103. Exothermic reactions absorb heat	BECAUSE	breaking covalent bonds always requires energy.
104. The solubility of gases in liquids does not depend upon pressure	BECAUSE	the vapor pressure of a substance is independent of external pressure.
105. MgO has a high melting point	BECAUSE	highly charged ions result in strong ionic forces and high lattice energies.
106. The ground state electron configuration orbitals of elemental Cu is $[\text{Ar}] 4s^1 3d^{10}$	BECAUSE	completely half-filled and filled <i>d</i> bestow special electronic stabilization.
107. Isotopes of a particular element have nearly identical chemical behavior	BECAUSE	they have identical electron configurations.
108. In an electrochemical cell, the electrode that is the site of reduction is called the anode	BECAUSE	oxidation always occurs at the cathode.
109. The addition of acid to a solution buffered to pH 7 slightly lowers the pH	BECAUSE	the addition of acids to any neutral solution always lowers the pH.
110. Saltwater boils at a higher temperature than pure water	BECAUSE	the presence of salt increases the vapor pressure of water.
111. BF_3 has a tetrahedral geometry	BECAUSE	the central B atom does not have a complete stable octet.
112. Hydrogen peroxide, H_2O_2 , is a good oxidizing agent	BECAUSE	the hydrogen in H_2O_2 has a +1 oxidation number.
113. Hydrogen gas (H_2) is considered a perfectly ideal gas	BECAUSE	hydrogen atoms interact with each other via hydrogen bonds.
114. Electrolysis of water requires the input of energy	BECAUSE	the products formed, H_2 and O_2 , possess more chemical potential energy than H_2O .
115. By mass, oxygen is the most abundant element in the human body	BECAUSE	it is principally found as O_2 in the bloodstream.

RETURN TO THE SECTION OF YOUR ANSWER SHEET YOU STARTED FOR **CHEMISTRY** AND ANSWER QUESTIONS 33–69.

Part C

Directions: Each of the questions or incomplete statements below is followed by five suggested answers or completions. Select the one that is best in each case and then fill in the corresponding oval on the answer sheet.

33. Choose the answer below that accurately describes the correct molecular shape for the molecule XeOF_4 .
- (A) Tetrahedral
 - (B) Trigonal pyramidal
 - (C) Trigonal bipyramidal
 - (D) Square pyramidal
 - (E) Flat
34. For the radioactive atom ^{99}Tc , what is the correct number of protons and neutrons?
- (A) 43 protons and 56 neutrons
 - (B) 43 protons and 99 neutrons
 - (C) 56 protons and 43 neutrons
 - (D) 56 protons and 99 neutrons
 - (E) Cannot be determined
35. Which one of the following acids is NOT strong?
- (A) HCl
 - (B) HBr
 - (C) HNO_3
 - (D) H_3PO_4
 - (E) H_2SO_4
36. Identify the equation used to determine the amount of heat required to melt 10 grams of ice.
- (A) $Q = mC_{sp}\Delta T$
 - (B) $Q = n\Delta H$
 - (C) $KE = \frac{1}{2}mv^2$
 - (D) $PE = mgh$
 - (E) $PV = nRT$
37. Identify the correct ground state electron configuration for Cr.
- (A) $[\text{Ar}] 3s^2 3d^4$

- (B) $[\text{Ar}] 3s^2 3d^5$
- (C) $[\text{Ar}] 4s^2 3d^5$
- (D) $[\text{Ar}] 4s^2 3d^4$
- (E) $[\text{Ar}] 4s^1 3d^5$

38. What is the hydroxide concentration for a solution with a pH of 10 at 25°C?
- (A) $10^{-14} M$
 - (B) $10^{-10} M$
 - (C) $10^{-7} M$
 - (D) $10^{-4} M$
 - (E) $10^{-1} M$
39. Five hundred milliliters of solution of 0.1 M NaBr has how many milligrams of bromine?
- (A) 200 mg
 - (B) 400 mg
 - (C) 2,000 mg
 - (D) 4,000 mg
 - (E) 20,000 mg
40. According to the ideal gas law, what is the approximate volume that will be occupied by 0.5 mole of an ideal gas at 30°C and 3 atm pressure (gas constant $R = 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$)?
- (A) Less than 1 L
 - (B) 5 L
 - (C) 10 L
 - (D) 15 L
 - (E) More than 20 L
41. Given that $\Delta G = \Delta H - T\Delta S$, how is the spontaneity of an endothermic reaction expected to change with decreasing T ?
- (A) Becomes less spontaneous
 - (B) Becomes more spontaneous
 - (C) Does not change
 - (D) Decreases at first but then increases
 - (E) Insufficient information to make a conclusion
42. Identify the element with the greatest first ionization energy.
- (A) Ce

- (B) C
- (C) Cl
- (D) Ca
- (E) Cs

43. Identify the molecule/ion with the greatest potential to act as a Lewis acid.

- (A) CH_3^+
- (B) CN^-
- (C) NH_3
- (D) BF_4^-
- (E) CO_2



44. Which coefficient balances the reaction given above?

- (A) 2
- (B) 4
- (C) 5
- (D) 6
- (E) 8

45. A 100-milliliter solution containing AgNO_3 was treated with excess NaCl to completely precipitate the silver as AgCl . If 5.7 g AgCl was obtained, what was the the concentration of Ag^+ in the original solution?

- (A) 0.03 M
- (B) 0.05 M
- (C) 0.12 M
- (D) 0.30 M
- (E) 0.40 M

46. Identify which of the following statements is FALSE.

- (A) The vapor pressure of a liquid decreases with increasing atmospheric pressure.
pressure.
- (B) The value of an equilibrium constant is dependent on temperature.
- (C) The rate of a spontaneous reaction cannot be determined solely by its Gibbs free
free energy.
- (D) During a phase transition, the temperature of a substance must be constant.
constant.
- (E) The addition of a catalyst to a reaction at equilibrium has no net effect on the
the system.

47. Which of the following compounds would be expected to have the greatest lattice binding energy?

- (A) LiNO_3
- (B) LiF
- (C) KI
- (D) NH_4Br
- (E) CsNO_3

48. The daughter nucleus formed when ^{18}F undergoes positron emission is

- (A) ^{14}N
- (B) ^{16}O
- (C) ^{18}O
- (D) ^{19}F
- (E) ^{20}Ne

49. Which of the following reactions produces a yellow precipitate?

- (A) $\text{NaOH}(aq) + \text{HCl}(aq) \rightarrow \text{NaCl}(s) + \text{H}_2\text{O}$
- (B) $\text{NaOH}(aq) + \text{BaCl}(aq) \rightarrow \text{BaOH}(s) + \text{NaCl}(aq)$
- (C) $\text{Pb}(\text{NO}_3)_2(aq) + 2\text{KI}(aq) \rightarrow 2\text{KNO}_3(aq) + \text{PbI}_2(s)$
- (D) $\text{CuO}(s) + \text{Mg}(s) \rightarrow \text{Cu}(s) + \text{MgO}(s)$
- (E) $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$



50. In the electrochemical cell described by the cell diagram above, what reaction occurs at the anode?

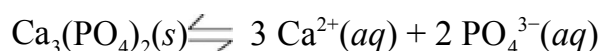
- (A) $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$
- (B) $\text{Zn}^{2+} + 2\text{e}^- \rightarrow \text{Zn}$
- (C) $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$
- (D) $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$
- (E) $\text{Zn} + \text{Cl}_2 \rightarrow \text{ZnCl}_2$

51. Given the reaction $\text{A} \rightarrow \text{B} + \text{C}$, where ΔH_{rxn} is negative, what effect would increasing the temperature (at constant pressure) have on the system at equilibrium?

- (A) No change
- (B) Cannot be determined
- (C) Shift to the right
- (D) Shift to the left for $K < 1$ and to the right for $K > 1$

52. An unknown acid solution was presumed to be either HCl or H₂SO₄. Which one of the following salt solutions would produce a precipitate when added to H₂SO₄ but not when added to HCl?

(A) LiNO₃
(B) NH₄NO₃
(C) CsNO₃
(D) Ba(NO₃)₂
(E) AgNO₃



53. What is the equilibrium expression for the dissolution of Ca₃(PO₄)₂ where the above is true?

(A) $K_{sp} = [\text{Ca}^{2+}]^3[\text{PO}_4^{3-}]^2$
(B) $K_{sp} = [\text{Ca}^{2+}]^2[\text{PO}_4^{3-}]^3$
(C) $K_{sp} = [\text{Ca}^{2+}][\text{PO}_4^{3-}]/[\text{Ca}_3(\text{PO}_4)_2]$
(D) $K_{sp} = [\text{Ca}^{2+}]^3[\text{PO}_4^{3-}]^2/[\text{Ca}_3(\text{PO}_4)_2]$
(E) $K_{sp} = [\text{Ca}^{2+}]^2[\text{PO}_4^{3-}]^3/[\text{Ca}_3(\text{PO}_4)_2]$

54. Which of the following represents a conjugate acid/base pair?

(A) Na⁺/Cl⁻
(B) HCl/H⁺
(C) H₂CO₃/CO₃²⁻
(D) NH₃/NH₄⁺
(E) K⁺/OH⁻

55. An unknown solution having a pH of 3.5 was titrated with 0.1 M NaOH. Analysis of the resulting titration curve showed a single equivalence point at pH 7. Therefore, which of the following could be the unknown solute in the initial solution?

(A) HF
(B) HCl
(C) LiOH
(D) NH₃
(E) H₂SO₄

56. Acid/base titration experiments could be used to determine all of the following directly EXCEPT

- (A) the acid concentration of an acidic solution
- (B) the alkalinity of a basic solution
- (C) the pK_a of an unknown weak acid
- (D) whether an unknown acid is monoprotic or polyprotic
- (E) the molecular weight of an unknown acid or base

57. What is the correct term for the phase change from gas directly to solid?

- (A) Deposition
- (B) Sublimation
- (C) Liquefaction
- (D) Fusion
- (E) Vaporization

58. What is the correct name for a straight-chained organic compound with the molecular formula C_3H_8 ?

- (A) Methane
- (B) Ethane
- (C) Methylethane
- (D) Propane
- (E) Isopropane

59. If the pH of a solution is changed from 1 to 3 with the addition of an antacid, what percentage of $[H^+]$ was neutralized?

- (A) 2%
- (B) 10%
- (C) 20%
- (D) 90%
- (E) 99%

60. Which of the following statements is the most accurate with regard to the significance of Avogadro's number, 6.02×10^{23} ?

- (A) It is the conversion factor between grams and atomic mass units.
- (B) It is a universal physical constant just as the speed of light.
- (C) It is the number of particles that is required to fill a 1-liter container.
- (D) It is the inverse diameter of an H atom.
- (E) It is the number of electrons in the universe.

Questions 61–64 refer to the following data at standard conditions.

	Appearance	Reactions with dilute HCl	Reaction with dilute HNO ₃
Unknown metal #1	Dull gray solid with white oxide coating	Dissolved with bubbles of clear gas	Dissolved with bubbles of clear gas
Unknown metal #2	Solid; lustrous, smooth silver-gray surface	No reaction	Dissolved with bubbles of orange gas

61. Unknown metal #1 could be

- (A) mercury
- (B) copper
- (C) zinc
- (D) iron
- (E) silver

62. Unknown metal #2 could be

- (A) carbon
- (B) copper
- (C) zinc
- (D) sodium
- (E) silver

63. The addition of dilute HCl to unknown metal #1 produced a transparent gas. What is the likely identity of this gas?

- (A) Cl₂
- (B) H₂
- (C) O₂
- (D) CO₂
- (E) NO₂

64. The addition of dilute HNO₃ to unknown metal #2 produced an orange gas. What is the likely identity of this gas?

- (A) Cl₂
- (B) H₂
- (C) O₂
- (D) CO₂
- (E) NO₂

65. Which of the following solutions is the product of the neutralization reaction between 10 ml 0.2 M KOH and 10 ml 0.2 M HI?
- (A) 0.1 M KI_3
 - (B) 0.1 M KI
 - (C) 0.2 M KI
 - (D) 0.4 M KI
 - (E) 0.4 M HOH
66. Which of the following is true regarding an Ne atom with a mass number of 20 and an O^{2-} ion with a mass number of 16?
- (A) They contain the same number of protons.
 - (B) They contain the same number of neutrons.
 - (C) They contain the same number of protons plus neutrons.
 - (D) They are isoelectronic.
 - (E) They are isomers.
67. Which of the following statements is NOT correct regarding chemical catalysts?
- (A) They are not consumed during the chemical reaction.
 - (B) They cannot make nonspontaneous reactions occur.
 - (C) They do not have to be the same phase as the reactant molecules.
 - (D) They shift equilibrated reactions to the product's side.
 - (E) Enzymes are biological catalysts.
68. Most elements are solids at 25°C and 1 atm pressure, the exception being the 11 elements that are gases and 2 that are liquids. What 2 elements are liquids?
- (A) Hg and Br
 - (B) Hg and I
 - (C) Ag and Kr
 - (D) Au and Kr
 - (E) Pt and Co
69. A student conducted an experiment and obtained three values during three repetitive trials: 1.65, 1.68, 1.71. Later, the student discovered that the true value was 2.37. In contrast to the real value, the experimental results should be characterized as
- (A) not accurate and not precise
 - (B) accurate but not precise
 - (C) not accurate but precise
 - (D) accurate and precise

STOP

**If you finish before time is called, you may check your work on this section only.
Do not turn to any other section in the test.**

HOW TO SCORE THE PRINCETON REVIEW PRACTICE SAT CHEMISTRY SUBJECT TEST

When you take the real exam, the proctors will collect your test booklet and bubble sheet and send your answer sheet to New Jersey, where a computer looks at the pattern of filled-in ovals on your answer sheet and gives you a score. We couldn't include even a small computer with this book, so we are providing this more primitive way of scoring your exam.

Determining Your Score

STEP 1

Using the answer key on the next page, determine how many questions you got right and how many you got wrong on the test. Remember: Questions that you do not answer don't count as either right or wrong answers.

STEP 2

List the number of right answers here.

(A) _____

STEP 3

List the number of wrong answers here. Now divide that number by 4. (Use a calculator if you're feeling particularly lazy.)

(B) _____ $\div$ 4 = (C) _____

STEP 4

Subtract the number of wrong answers divided by 4 from the number of correct answers. Round this score to the nearest whole number. This is your raw score.

A) _____ $-$ (C) _____ = _____

STEP 5

To determine your real score, take the number from Step 4 above, and look it up in the left column of the Score Conversion Table on [this page](#); the corresponding score on the right is your score on the exam.

ANSWERS TO THE PRINCETON REVIEW PRACTICE SAT CHEMISTRY SUBJECT TEST 3

Question Number	Correct Answer	Right	Wrong	Question Number	Correct Answer	Right	Wrong
1.	E	_____	_____	46.	A	_____	_____
2.	A	_____	_____	47.	B	_____	_____
3.	B	_____	_____	48.	C	_____	_____
4.	B	_____	_____	49.	C	_____	_____
5.	C	_____	_____	50.	A	_____	_____
6.	A	_____	_____	51.	E	_____	_____
7.	E	_____	_____	52.	D	_____	_____
8.	D	_____	_____	53.	A	_____	_____
9.	B	_____	_____	54.	D	_____	_____
10.	D	_____	_____	55.	B	_____	_____
11.	C	_____	_____	56.	E	_____	_____
12.	B	_____	_____	57.	A	_____	_____
13.	A	_____	_____	58.	D	_____	_____
14.	B	_____	_____	59.	E	_____	_____
15.	C	_____	_____	60.	A	_____	_____
16.	D	_____	_____	61.	C	_____	_____
17.	E	_____	_____	62.	E	_____	_____
18.	D	_____	_____	63.	B	_____	_____
19.	A	_____	_____	64.	E	_____	_____
20.	B	_____	_____	65.	B	_____	_____
21.	C	_____	_____	66.	D	_____	_____
22.	E	_____	_____	67.	D	_____	_____
23.	D	_____	_____	68.	A	_____	_____
24.	B	_____	_____	69.	C	_____	_____
25.	B	_____	_____				
26.	A	_____	_____				
27.	C	_____	_____				
28.	C	_____	_____	101.	T, T, CE		
29.	D	_____	_____	102.	T, F		
30.	A	_____	_____	103.	F, T		
31.	A	_____	_____	104.	F, T		
32.	E	_____	_____	105.	T, T, CE		
33.	D	_____	_____	106.	T, T, CE		
34.	A	_____	_____	107.	T, T, CE		
35.	D	_____	_____	108.	F, F		
36.	B	_____	_____	109.	T, T, CE		
37.	E	_____	_____	110.	T, F		
38.	D	_____	_____	111.	F, T		
39.	D	_____	_____	112.	T, T		
40.	B	_____	_____	113.	F, F		
41.	E	_____	_____	114.	T, T, CE		
42.	B	_____	_____	115.	T, F		
43.	A	_____	_____	116.	T, T		
44.	D	_____	_____				
45.	E	_____	_____				

THE PRINCETON REVIEW PRACTICE SAT CHEMISTRY SUBJECT TEST SCORE CONVERSION TABLE

Raw Score	Scaled Score	Raw Score	Scaled Score	Raw Score	Scaled Score
85	800	45	620	5	390
84	800	44	620	4	390
83	800	43	610	3	380
82	800	42	610	2	380
81	800	41	600	1	370
80	800	40	590	0	370
79	800	39	590	-1	370
78	790	38	580	-2	360
77	780	37	580	-3	360
76	780	36	570	-4	350
75	780	35	560	-5	340
74	780	34	560	-6	340
73	780	33	550	-7	330
72	770	32	550	-8	330
71	770	31	540	-9	320
70	750	30	530	-10	310
69	750	29	530	-11	310
68	740	28	520	-12	300
67	740	27	520	-13	300
66	740	26	520	-14	290
65	730	25	510	-15	280
64	730	24	510	-16	280
63	710	23	500	-17	270
62	710	22	500	-18	270
61	710	21	490	-19	260
60	700	20	480	-20	250
59	700	19	480	-21	250
58	690	18	470		
57	690	17	470		
56	680	16	460		
55	680	15	450		
54	680	14	450		
53	670	13	440		
52	670	12	440		
51	660	11	430		
50	650	10	420		
49	650	9	420		
48	630	8	410		
47	630	7	410		
46	630	6	400		