Steric Number*	# of Atoms Bonded to Central Atom	# of Lone Pairs on Central Atom	Examples	Electron Pair Geometry (Hybridization)	Shape/Geometry of Atoms (VESPER class)	Bond Angles
1	1	0	H_2	Linear (s)	Linear (AB)	180°
2	2	0	BeH_2, $BeCl_2$, CO_2, CS2, HCN	Linear (sp)	Linear (AB_2)	180°
	1	1	N_2, CO, CN^-		Linear (AB_2)	180°
3	3	0	BCl_3, BF_3, $AlCl_3$, CH_2O, SO3, CO_3^{2-}	Trigonal Planar (sp^2)	Trigonal Planar (AB_3)	120°
	2	1	O_3, $SnCl_2$, SO_2, NO_2^-		Bent/V-shaped (AB_2E)	118 ° (<120°)
	1	2	O_2		Linear (ABE_2)	120°
4	4	0	CH_4, $SiCl_4$, $POCl_3$, SO_4^{2-}	Tetrahedral (sp^3)	Tetrahedral (AB_4)	109.5°
	3	1	NH_3, PF_3, PH_3, H_3O^+		Trigonal Pyramidal (AB_2E)	107.5° (<109°)
	2	2	H_2O, H_2S, $SeBr_2$, BO_2^-, ICl_2^+		Bent/V-shaped (AB_2E_2)	104.5° (<<109°)
	1	3	F_2, Cl_2, HF, OH^-		Linear (ABE_3)	180°
5	5	0	PF_5, PCl_5, PH_5, AsF_5, SbF_5, $IO_3F_2^-$	Trigonal Bipyramidal (sp^3d)	Trigonal Bipyramidal (AB_5)	90° and 120°
	4	1	SF_4, $TeCl_4$, SeH_4, IF_4^+		Seesaw/Sawhorse (AB_4E)	<90° and <120°
	3	2	ICl_3, ClF_3, BrF_3		T-shaped (AB_3E_2)	<90°
	2	3	XeF_2, BrF_2^-, I_3^-, ICl_2^-		Linear (AB_2E_3)	180°
6	6	0	SF_6, $SeCl_6$, PF_6^-, SiF_6^{2-}	Octahedral (sp^3d^2)	Octahedral (AB_6)	90°
	5	1	BrF_5, IF_5, $XeOF_4$, $SbCl_5^{2-}$		Square Pyramidal (AB_5E)	<90°
	4	2	XeF_4, ICl_4^-, BrF_4^-		Square Planar (AB_4E_2)	90°

*Steric Number AKA Electron Domain: total number of atoms and lone pairs connected to central atom

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